**\*\*\*Memorize by first exam\*\*\***

* Diatomic Elements (H2 N2 O2 F2 Cl2 Br2 I2)
* “-ate” polyatomic ions (top table on green sheet)
* Ways to modify polyatomic ions/names (-ite, per-, hypo-, hydrogen-, dihydrogen-)
* Additional Common Polyatomic Ions

|  |  |
| --- | --- |
| **Name** | **Chemical Formula** |
| Hydroxide | OH- |
| Cyanide | CN- |
| Acetate | C2H3O2- |
| Ammonium | NH4+ |

* Metric Prefixes (nano, micro, milli, centi, kilo)
* Conversions/Constants
* 1 cm3 = 1mL (exactly) 1in = 2.54 cm(exactly) DH2O = 1.0 g/cm3
* Equations

**\*\*\*Memorize by second exam\*\*\***

* VSEPR Shapes, bond angles
* Avogadro’s # = 6.022 x1023
* Equations

**\*\*\*Memorize by third exam\*\*\*** *I (seth) reserve the right to add to this list*

* Strong Acids (HCl, HNO3, H2SO4)
* Equations

**\*\*\*Memorize by fourth exam\*\*\*** *I (seth) reserve the right to add to this list*

* Types of nuclear emissions/decay (alpha, beta, gamma)